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Assessment Reaction Rates  
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# Chapter Assessment Reaction Rates Answers

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### Answers

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Chapter 6. Big Idea #4: Chemical Reactions and their Rates. CHAPTER 6 ANSWERS AND EXPLANATIONS. Multiple-Choice. 1. D The answer to (A) shows all of the reactants and products that are part of this reaction. However, species that appear on both sides need to cancel. That eliminates B, one C, and D, leaving  $2A + C \rightarrow E$ . 2.

#### **CHAPTER 6 ANSWERS AND EXPLANATIONS - Chemical Reactions ...**

Reaction Rates in Analysis: Test Strips for Urinalysis Physicians often use disposable test strips to measure the amounts of various substances in a patient's urine ( Figure 3 ). These test strips contain various chemical reagents, embedded in small pads at various locations along the strip, which undergo changes in color upon exposure to

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sufficient concentrations of specific substances.

### **12.1 Chemical Reaction Rates - Chemistry**

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a state of balance in which the rates of the forward and reverse reactions are equal; no net change in the amount of reactants and products occurs in the chemical system (18.2) equilibrium position the relative concentrations of reactants and products of a reaction that has reached equilibrium; indicates whether the reactants or products are favored in the reversible reaction (18.2)

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(f) decreases rate of reaction (g) increases rate of reaction (h) decreases rate of reaction (i) increases rate of



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### Answers

reaction (j) increases rate of reaction 2.

Applying Knowledge Four factors

affecting the rate of reactions Page 118

1. (a) line Y (b) line X (c) line Y (d) line X

(e) line Y (f) line X (g) line Y (h) line X 2.

(a) surface area (b) catalyst (c)

temperature (d) concentration

### **Section 6.2 Factors Affecting the Rate of Chemical Reactions**

D)  $k$  is the reaction rate constant E) The reaction is second order overall. 15) The half-life of a first-order reaction \_\_\_\_\_.

A) is constant B) is the time necessary for the reactant concentration to drop to half its original value C) can be calculated from the reaction rate constant D) does not depend on the initial reactant concentration

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560 Chapter 16 • Reaction Rates Section

116.16.1 A Model for Reaction Rates

MAIN Idea Collision theory is the key to understanding why some reactions are

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faster than others. Real-World Reading Link Which is faster: walking to school, or riding in a bus

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Question 2 The concentration of a reactant is measured at two time intervals as a reaction proceeds. At the first time interval, the concentration of the reactant is 0.5 mol L<sup>-1</sup>; 20 seconds later, the concentration is 0.45 mol L<sup>-1</sup>. At what rate is the reactant being consumed?

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